**Task 5 Electrochemical Cell Part 2 Activity Analysis Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Task description:** This session is for you to use the data acquired in **Part 1** answer the following questions:

**TOTAL: \_\_\_\_\_\_ / 27 Marks**

**Question 1:**



**Cell 1 (Cu/Cu2+ Zn/Zn2+): (3 marks)**

**Cell 2 (Cu/Cu2+  Pb/Pb2+): (3 marks)**

**Cell 3 (Zn/Zn2+ Pb/Pb2+): (3 marks)**

**Question 2: How do the voltages from the experiment compare to the theoretical values derive using your data sheet? Account for the difference. (3 marks)**

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**Question 3: Account for your observations when Na2S solution was added to the Cu(NO3)2 solution in the Zn/Zn2+ Cu/Cu2+ cell. (2 marks)**

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**Question 4: Use your data from the experiment to arrange the metals in decreasing oxidising strength. Show and explain how you determined your rank. (2 marks)**

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**Question 5: What is the independent variable?**  **(1 mark)**

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**Question 6: What is the dependent variable? (1 mark)**

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**Question 7: List three controlled variables. (1 mark)**

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**Question 8: List two systematic errors associated with the activity (2 marks)**

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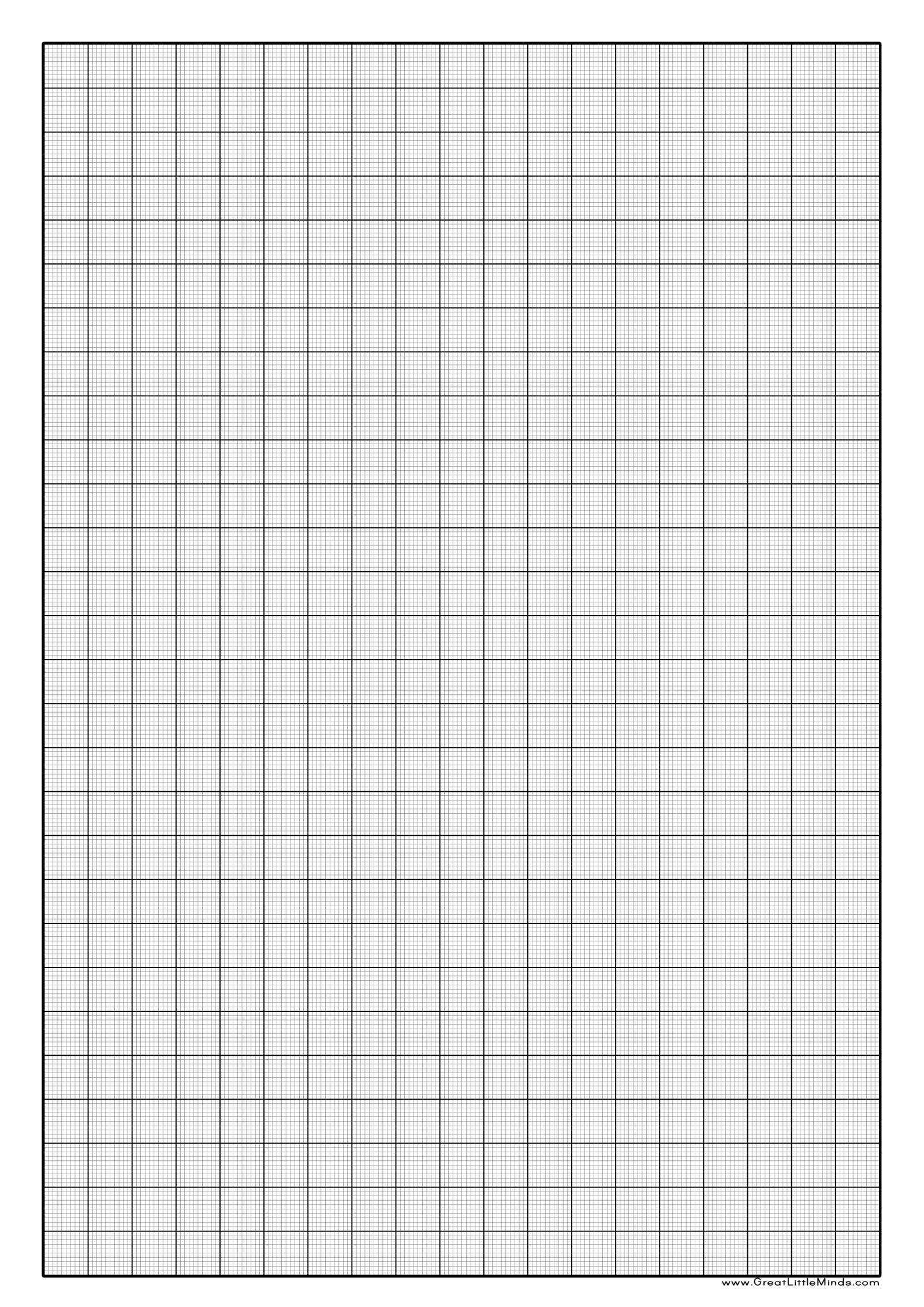
**Question 9: Considering the data explain which type of graph you should use.**

**(2 marks)**

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**Question 10: Graph your data on the graph paper on the next page. If you have no data, use the theoretical voltages. (4 marks)**



**END OF PART 2**